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~~Atom~~ 4-1a Intro

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Atomic Orbitals,
and Electron
Configurations*

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~~Configuration~~

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Elements in

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Arrangement in

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of Atom | SPM

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Distribution of

Electrons |

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Atom | How
Arrangement Of
Electrons
distributed |

Class 9 **Energy
Levels, shells,
SubLevels \u0026
Orbitals** *How*

*does the
electron move
around the atom?*

How to write
electron
configurations

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and what they
are Quantum
Mechanics Part 3
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Electrons and

Sublevels How

Small Is An

Atom? Spoiler:

Very Small.

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Model of an Atom
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configuration of
atoms using

Aufbau, Pauli's
principle and

Hund's rule -

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levels,
sublevels,
& orbitals
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Electrons~~

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9 chemistry

chapter 4

Arrangements of
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~~Electron~~

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Valence

Electrons and

the Periodic

TableChapter 3

Arrangement of

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(Section 3.6)

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Understanding the Atom_OLD

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Atoms. 33 terms.

Chem Chapter 4.

25 terms.

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terms. Macbeth
Acts 3, 4, 5. 8

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the Electrons
Chapter 4
(Electron
Configurations)
Electron
Behavior. ...

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Arrangement Of
Electrons

arrangement by
wavelength or
frequency for
all forms of
electromagnetic
radiation. Parts
of the wave. Wav
elength- λ
(?) The distance
between
corresponding
points on
adjacent waves.

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Chapter 4
Units: m, nm,
cm, or Å
Arrangement Of
Electrons

Arrangement of
the Electrons

Chapter 4

CHAPTER 4 REVIEW

Arrangement of
Electrons in

Atoms SECTION 3

SHORT ANSWER

Answer the

following

questions in the

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space provided.

1. State the Pauli exclusion principle, and use it to explain why electrons in the same orbital must have opposite spin states. The Pauli exclusion principle states that no two

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atom may have
the
Arrangement Of
Electrons

4 Arrangement of
Electrons in
Atoms

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Electrons in
Atoms Section
4-1: The
Development of a
New Atomic Model

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Regular

Schedule: with

lab(s): 3 days

without lab(s):

2 days Block

Schedule: with

lab(s): 1 1/2

days without

lab(s): 1 day

Objectives 1.

Explain the

mathematical

relationship

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speed,
wavelength, and
frequency of ...

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atoms Taken from
the book Modern
Chemistry by
Holt, Rinehart,
and Winston on
Chapters 4 and
5, which deals
with electrons
and the periodic
table. Includes
the chapter
vocabulary and a
few other useful

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of electrons in
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Review
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[Arrangement of
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Chapter 5 [The
Arrangement Of
Periodic Law]

Chapter 6

[Chemical
Bonding] ...

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Electrons.

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Spectra .

Absorption and
Emission spectra
for the elements

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• Bohr model of
the atom .
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Dalton's atomic
theory quiz.

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the Final From
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the arrangement
of electrons in
an atom. Ground-
State Electron
Configuration.
the lowest-
energy
arrangement of

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the electrons
for each
Arrangement Of
element. Aufbau
Principle
Electrons

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Vocabulary:
Arrangement of
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Atoms ...
Elements &
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Configurations
Elements of the

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6th and 7th
periods contain
“f” orbitals. Do
Electrons

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Modern Chemistry
– Chapter 4:
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Electrons in

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Atoms.
Electromagnetic
Radiation.

Electromagnetic
Spectrum.

Wavelength.

Frequency. The
radiation
associated
within electric
and magnetic
field; i... All
of the

frequencies or

Where To Download

wavelengths of
electromagnetic
radiation...
Electrons

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chemistry
arrangement
electrons modern
...

Modern Chemistry
29 Arrangement
of Electrons in
Atoms CHAPTER 4
REVIEW

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Arrangement of
Electrons in
Atoms SECTION 3
SHORT ANSWER

Answer the
following
questions in the
space provided.

1. State the
Pauli exclusion
principle, and
use it to
explain why
electrons in the

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Chapter 4
Arrangement Of
Electrons
same orbital
must have
opposite spin
states.

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Electrons in
Atoms Section
4-3: Electron

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1 day Objectives

1. List the

total number of

electrons needed

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to fully occupy
each main energy
level. 2.

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CHAPTER 4:
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ELECTRONS IN
ATOMS The
following pages
contain the bulk

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for the chapter
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Electrons
orbitals of

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equal energy are each occupied by one electron before any orbital is occupied by a second electron, and all electrons in singly occupied orbitals must have the same spin Pauli's exclusion Page

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5/26 Chapter 4

Arrangement Of

Chapter 4

Electrons

Arrangement Of

Electrons -

Bespokify

The Pauli

exclusion

principle states

that no two

electrons in an

atom may have

the Chapter Four

[Arrangement of

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Atoms] CHEMISTRY
Arrangement Of
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lowest energy
state of an atom
is its ground
state.

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Rutherford's

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Did not explain

how the

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negatively
charged
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electrons
distributed

itself in the
electron cloud
around the
positively
charged nucleus
without being
attracted to
each other Holt
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